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Key indicators

Single-crystal X-ray study T = 180 K Mean σ (C–C) = 0.002 Å R factor = 0.034 wR factor = 0.089 Data-to-parameter ratio = 11.7

For details of how these key indicators were automatically derived from the article, see http://journals.iucr.org/e.

N,N-Bis[2-(2-pyridyl)ethyl]hydroxylamine

The crystal structure of the title compound, $C_{14}H_{17}N_3O$, at 180 K contains intermolecular hydrogen bonds between the hydroxyl group and the N atom of one pyridyl ring.

Comment

The title compound, (I) (Fig. 1), has recently appeared in the literature as a convenient source of the bis(2-pyridylethyl)amino skeleton, commonly incorporated in ligands employed for biomimetic coordination chemistry (Leaver *et al.*, 2003). In its crystal structure at 180 K, both pyridyl rings adopt anticlinal conformations with respect to the hydroxyl group so that the C_5H_4N and OH groups form angles of 120° with respect to each other in a Newman-type projection along the $(CH_2)_2N(CH_2)_2$ core of the molecule (Fig. 2). A dihedral angle of 67.5 (1)° exists between the least-squares planes of the two pyridyl rings. Intermolecular hydrogen bonds are formed between the hydroxyl group and the N atom of one pyridyl ring $[H1\cdots N3^i = 1.94 (3) \text{ Å}, O1-H1\cdots N3^i = 164 (2)^\circ$; symmetry code (i): 1 + x, y, z].



Experimental

The title compound was prepared according to the synthetic procedure of Leaver *et al.* (2003).

C ₁₄ H ₁₇ N ₃ O	Mo $K\alpha$ radiation
$M_r = 243.31$	Cell parameters from 3638
Orthorhombic, $P2_12_12_1$	reflections
a = 5.8501 (4) Å	$\theta = 2.7 - 28.4^{\circ}$
b = 14.7210(9) Å	$\mu = 0.08 \text{ mm}^{-1}$
c = 15.2578 (10) Å	T = 180 (2) K
$V = 1313.99 (15) \text{ Å}^3$	Block, colourless
Z = 4	$0.40 \times 0.30 \times 0.10 \text{ mm}$
$D_{\rm r} = 1.230 {\rm Mg} {\rm m}^{-3}$	



The molecular structure, showing displacement ellipsoids at the 50% probability level. H atoms bound to C are shown as spheres of arbitrary radius.

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organic papers

Data collection

Bruker-Nonius X8APEX-II CCD 1958 independent reflections diffractometer 1681 reflections with $I > 2\sigma(I)$ Thin-slice ω and φ scans $R_{\rm int}=0.026$ $\theta_{\rm max} = 28.7^\circ$ Absorption correction: multi-scan (SADABS; Sheldrick, 2003) $h = -7 \rightarrow 7$ $k = -17 \rightarrow 19$ $T_{\rm min}=0.885,\ T_{\rm max}=0.992$ 9849 measured reflections $l = -20 \rightarrow 19$ Refinement Refinement on F^2 $w = 1/[\sigma^2(F_o^2) + (0.0492P)^2]$ $R[F^2 > 2\sigma(F^2)] = 0.034$ $wR(F^2) = 0.089$ where $P = (F_o^2 + 2F_c^2)/3$ $(\Delta/\sigma)_{\rm max} < 0.001$

S = 1.061958 reflections 167 parameters H atoms treated by a mixture of independent and constrained refinement

Table 1

Hydrogen-bonding geometry (Å, °).

$D - H \cdots A$	D-H	$H \cdots A$	$D \cdots A$	$D - \mathbf{H} \cdots A$
$O1 - H1 \cdots N3^i$	0.90 (3)	1.94 (3)	2.8174 (18)	164 (2)
Symmetry code: (i)	$1 \pm r = 7$			

+ 0.1214P]

 $\Delta \rho_{\rm max} = 0.23 \text{ e } \text{\AA}^{-3}$

 $\Delta \rho_{\rm min} = -0.15 \text{ e} \text{ \AA}^{-3}$

Symmetry code: (i) 1 + x, y, z

H atoms bound to C were positioned geometrically and allowed to ride during subsequent refinement [C-H = 0.95 Å, $U_{iso}(H)$ = $1.2U_{eq}(C)$ for H atoms bound to the pyridyl rings; C-H = 0.99 Å, $U_{\rm iso}({\rm H}) = 1.2U_{\rm eq}({\rm C})$ for H atoms of the methylene groups]. H1, associated with the hydroxyl group, was located in a difference Fourier map and refined freely with an isotropic displacement parameter. The absolute structure could not be determined and 1286 Friedel opposites were merged for the final cycles of refinement.

Data collection: APEX2 (Bruker-Nonius, 2003); cell refinement: SAINT (Bruker, 2003); data reduction: SAINT; program(s) used to solve structure: SHELXTL (Sheldrick, 2000); program(s) used to refine structure: SHELXTL; molecular graphics: SHELXTL; software used to prepare material for publication: SHELXTL.





Newman-type projection along the (CH₂)₂N(CH₂)₂ core of the molecule, showing the anticlinal conformations of the pyridyl rings with respect to the hydroxyl group.

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